

## *Engineering Chemistry 1st Year Corrosion Notes*







### **Engineering Chemistry 1st Year Corrosion**

Engineering Chemistry Tuesday, 28 February 2017. Unit - 3: Corrosion Science (BE 1st Year 2nd Semester, 2016-2017) Corrosion - Introduction - Definition - Causes & Effects of Corrosion - Types of Corrosion - (i) Dry / Direct Chemical Corrosion. Pilling-Bedworth Rule and its significance (ii) Wet or Electrochemical corrosion, Differences of ...

### **Engineering Chemistry : Unit - 3: Corrosion Science (BE ...**

Thus, greater is its tendency to become anodic and hence greater is the rate of corrosion. Thus, extent of corrosion depends upon the position of the metal in the galvanic series. When two metals are in electrical contact, greater is the difference in their positions in the electrochemical series, faster is the corrosion of anodic metal.

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### **Corrosion**

engineering chemistry 1st year corrosion notes 634538B1671101F6985631BCF9DC7F88 shares a common first year with the Mechanical Engineering Technology programs.

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anhydrous inorganic liquid, etc. The chemical corrosion is defined as the direct chemical attack of metals by the atmospheric gases present in the environment. Example: (i) Silver materials undergo chemical corrosion by Atmospheric H<sub>2</sub>S gas . (ii) Iron metal undergo chemical corrosion by HCl gas.

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