

Chapter 9 Review Stoichiometry Answers Section 3



Chapter 9 Review Stoichiometry Answers

CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed?

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ N_2 ; 2.0 mol a.

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Modern Chemistry 73 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. ____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

Chapter 9: Standard Review Worksheet

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Modern chemistry chapter 9 3 review stoichiometry answers

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on

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Answer Key Mole/Stoichiometry.Test.Review 1. 6.022×10^{23} particles ((atoms,(molecules)) (2. 1mole(= 6.022×10^{23} particles((1mole=molar(mass(1mole=22.4L(3. Calculate(the ...

Answer Key Mole/Stoichiometry.Test.Review

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